

Eindexamen scheikunde 1-2 vwo 2009 - I

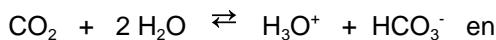
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Biobrandstofcel

- 1 Uit de diagrammen valt af te lezen hoe groot de concentratieveranderingen zijn van het glucose en de Fe^{2+} -ionen (in mmol L^{-1}) na 150 uur :

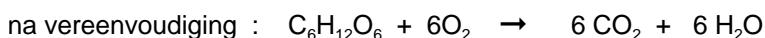
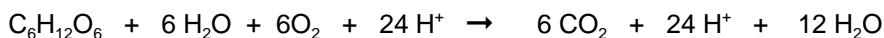
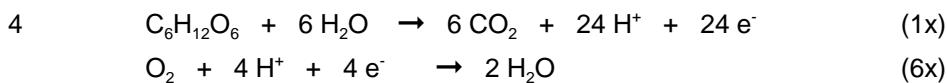
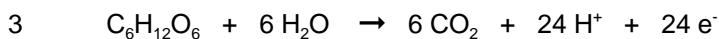
- afname glucose : $1,03 - 0,59 = 0,44 \text{ mmol L}^{-1}$
- toename Fe^{2+} : $12,6 - 2,0 = 10,6 \text{ mmol L}^{-1}$
- er reageert 0,44 mmol glucose met 10,6 mmol Fe^{3+}
dus 1,0 mol glucose reageert met $(10,6 / 0,44) = 24 \text{ mol } \text{Fe}^{3+}$
1 Fe^{3+} ion neemt één elektron op, 1 glucose-molecuul moet dus 24 electronen afstaan.

2 $\text{pH} = 6,90 \rightarrow [\text{H}_3\text{O}^+] = 1,26 \times 10^{-7} \text{ mol L}^{-1}$

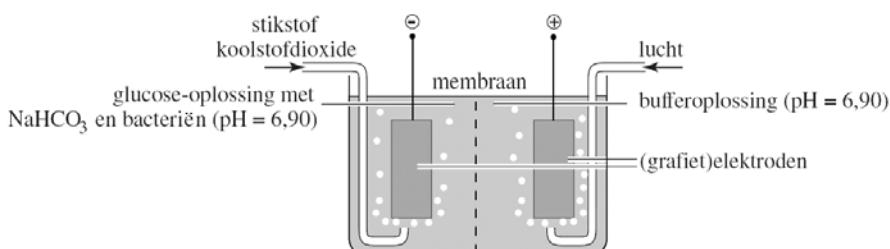


$$K_z = [\text{H}_3\text{O}^+] \times [\text{HCO}_3^-] / [\text{CO}_2] = 4,5 \times 10^{-7} = 1,26 \times 10^{-7} \times [\text{HCO}_3^-] / 0,011 \\ \rightarrow [\text{HCO}_3^-] = 0,0393 \text{ mol L}^{-1}$$

per L moet opgelost worden : $0,0393 \times 84,01 = 3,3 \text{ g NaHCO}_3$



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- 6
- $i = 0,20 \times 10^{-3} \text{ A} = 0,20 \times 10^{-3} \text{ C s}^{-1}$ en $t = 600 \text{ uur} = 600 \times 60 \times 60 = 2,16 \times 10^6 \text{ s}$
doorgestroomd : $0,20 \times 10^{-3} \times 2,16 \times 10^6 = 432 \text{ C}$
 - dat is : $432 / 9,65 \times 10^4 = 4,48 \times 10^{-3} \text{ mol e}^-$
 - daarvoor werd omgezet : $4,48 \times 10^{-3} / 24 = 1,86 \times 10^{-4} \text{ mol glucose}$
 - 83% van de glucose wordt gebruikt voor stroomlevering, dus nodig :
 $(100 / 83) \times 1,86 \times 10^{-4} = 2,25 \times 10^{-4} \text{ mol glucose}$
 - dat is : $2,25 \times 10^{-4} \times 180 = 4,0 \times 10^{-2} \text{ g glucose}$