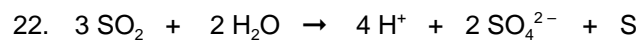
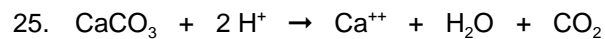


Kratermeer



23. $[\text{H}^+] = 10^{-0,2} = 0,6 \text{ mol/L}$

24. - 3 mol SO_2 reageert tot 1 mol S
- $90 \text{ ton SO}_2 = 90 \times 10^6 \text{ g SO}_2 = \{ 90 \times 10^6 / (32,06 + 32,00) \} = 1,4 \times 10^6 \text{ mol SO}_2$
- $1,4 \times 10^6 \text{ mol SO}_2$ reageert tot $1,4 \times 10^6 / 3 = 4,7 \times 10^5 \text{ mol S}$
- dat is : $4,7 \times 10^5 \times 32,06 = 1,5 \times 10^7 \text{ g S} = 15 \text{ ton S}$



27. - $14,75 \text{ mL } 0,011 \text{ M I}_2$ -opl bevat : $14,75 \times 0,011 = 0,16 \text{ mmol I}_2$
- $0,16 \text{ mmol I}_2$ reageert met $0,16 \text{ mmol SO}_3^{2-}$ dat is $1,6 \times 10^{-4} \text{ mol SO}_3^{2-}$

28. - $1,6 \times 10^{-4} \text{ mol SO}_3^{2-}$ is ontstaan uit $1,6 \times 10^{-4} \text{ mol SO}_2$
- dat is : $1,6 \times 10^{-4} \times (32,06 + 32,00) = 1,0 \times 10^{-2} \text{ g SO}_2$ per $0,30 \text{ dm}^3$ vulkaangas
- in $1,0 \text{ dm}^3$ vulkaangas zit : $(1,0 \times 10^{-2}) / 0,30 = 3,4 \times 10^{-2} \text{ g SO}_2$